



Centre for Academic Learner Services

Academic Upgrading

Chemistry 10 Placement Test

Study Guide

Updated: June 2015

Important Information about this Study Guide and the Placement Test

This study guide is designed to prepare students for the Academic Upgrading Chemistry 10 Placement test. An answer key is included at the end of this guide.

The test is taken in the Testing Centre (MC 221; see www.sait.ca/testing for hours, location and to book your test appointment). Contact Academic Upgrading to receive instructions for booking your testing appointment by emailing upgrading@sait.ca.

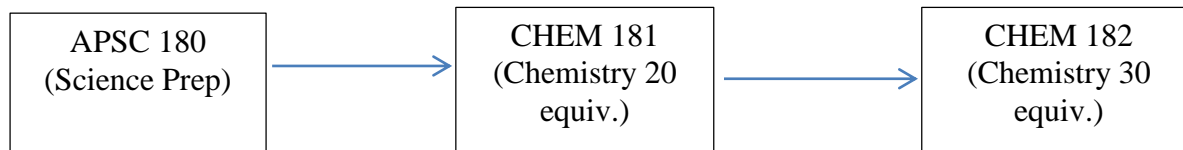
It is designed for Academic Upgrading placement purposes only. **This test may not be used for admission to any SAIT program; that is, this is not a SAIT admission exam. In addition, the results cannot be used at any other educational institution.**

The time allotted for the Chemistry 10 Placement test is 1 hour. The test consists of 20 questions and covers the chemistry material from Science 10. A mark of 65% is required to pass and allows entrance into CHEM 181.

Note: CHEM 181 is equivalent to Chemistry 20.

CHEM 181 is accepted as an admission requirement at other post-secondary institutions in Alberta, but you should always check with the post-secondary institution you are interested in attending (if it is not SAIT) to confirm it will accept the course.

SAIT Academic Upgrading Course Sequence



Introduction: Chemistry 10 Study Guide

- Review the practice exercises. You may use the solubility table and the periodic table of elements as provided at the end of the guide.
- Check your answers with the answer key provided at the end of this guide.
- You may choose to utilize a Science 10 Study Guide from the Calgary Public Library or bookstore for extra review. Make sure you focus on the chemistry chapters from such guides.
- You will be given a periodic table and solubility table during the test.
- You will need and are required to bring a scientific calculator for the test (graphing calculators are not permitted).

Multiple Choice

- John Dalton proposed four ideas in his model of the atom. Which of the following is **not** part of his model?
 - all matter is made of small indivisible particles
 - atoms of different elements have different properties
 - atoms are never created or destroyed during a chemical reaction
 - all the atoms of an element are identical in properties such as size and mass
- A certain chemical family is composed of elements that are soft, shiny, very reactive with water, and form ions with a charge of $1+$. This family could be
 - the halogens
 - the noble gases
 - the alkali metals
 - the alkaline-earth metals
- Which column in the periodic table contains elements with one electron in their valence energy level?
 - first on the left
 - first on the right
 - second from the left
 - second from the right
- Which two particles are approximately equal in mass?
 - proton and neutron
 - proton and electron
 - neutron and electron
 - none of these
- The magnesium ion, Mg^{+2} , has
 - 10 electrons and 10 protons
 - 10 electrons and 12 protons
 - 12 electrons and 10 protons
 - 12 electrons and 12 protons
- An atom of fluorine has 9 protons, 10 neutrons, and 9 electrons. Its mass number is
 - 9
 - 10
 - 18
 - 19

7. What is the formula for sodium carbonate?

- a) $\text{S}_2\text{CO}_{3(s)}$
- b) $\text{NaCO}_{(s)}$
- c) $\text{Na}_2\text{CO}_{3(s)}$
- d) $\text{Na}_3\text{CO}_{3(s)}$

8. What is the formula for aluminum hydroxide?

- a) $\text{AlOH}_{3(s)}$
- b) $\text{Al}_3\text{OH}_{(s)}$
- c) $\text{Al}(\text{OH})_{3(s)}$
- d) $\text{Al}(\text{III})\text{OH}_{(s)}$

9. Which of the following is an ionic compound?

- a) $\text{HCl}_{(aq)}$
- b) $\text{KCl}_{(s)}$
- c) $\text{ClO}_{3(g)}$
- d) $\text{NCl}_{3(g)}$

10. Which of the following properties are characteristic of an ionic compound?

- I It is malleable.
- II It is solid at room temperature.
- III Its solution conducts electricity.

- a) I and II only
- b) I and III only
- c) II and III only
- d) I, II, and III

11. Which of the following are very soluble?

- I Na_2S
- II CuBr
- III $\text{Sr}(\text{OH})_2$

- a) I and II only
- b) I and III only
- c) II and III only
- d) I, II, and III

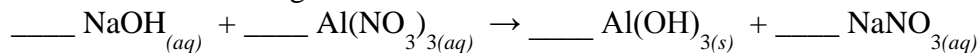
12. Which of the following is a general property of bases?

- a) taste sour
- b) turn litmus red
- c) conduct electricity
- d) react with Mg to produce hydrogen bubbles

13. Which of the following is an acid?

- a) $\text{CH}_4(g)$
- b) $\text{K}_3\text{PO}_4(aq)$
- c) $\text{H}_3\text{PO}_4(aq)$
- d) $\text{NaOH}(aq)$

14. Consider the following reaction:



The coefficient for $\text{Al(NO}_3)_3$ when the above equation is balanced is

- a) 1
- b) 2
- c) 3
- d) 4

15. The following reaction takes place when gasoline reacts with air:



This reaction is

- a) single replacement reaction
- b) double replacement reaction
- c) hydrocarbon combustion reaction
- d) decomposition reaction

Section II. Skills

Name or give the formula for each compound in questions 24 to 33. (1 mark each)

16. $\text{CaBr}_{2(s)}$

17. $\text{Au}_3\text{PO}_4(s)$

18. $\text{N}_2\text{O}_{4(g)}$

19. $\text{NH}_3(s)$

20. $\text{H}_2\text{SO}_4(aq)$

21. lead(IV) sulfide

22. methane

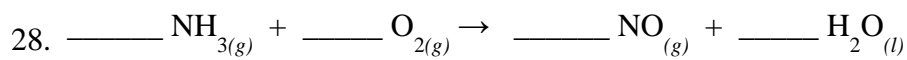
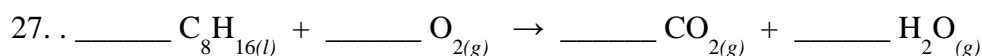
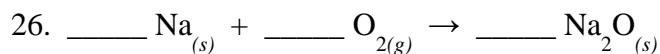
23. sulfur trioxide _____

24. hydrochloric acid _____

25. iron(II) nitride _____

Section III. Written Response

Balance the equations in questions 31, 32, and 33.



Predict the formulas of the products for each reaction below and WRITE them in the spaces provided, but do NOT balance the equations. Also, state the FULL name of the reaction type for each reaction.

	Products	Reaction Type
29. $\text{Na}_{(s)} + \text{Br}_{2(l)} \rightarrow$	_____	_____
30. $\text{CH}_{4(g)} + \text{O}_{2(g)} \rightarrow$	_____	_____
31. $\text{Al}_{(s)} + \text{CuCl}_{2(aq)} \rightarrow$	_____	_____
32. $\text{NaI}_{(aq)} + \text{Pb}(\text{NO}_3)_{2(aq)} \rightarrow$	_____	_____

Write balanced formula equations for the reactions in questions 33 and 34.

33. Aqueous ammonium sulfide and aqueous lead(II) nitrate are mixed together. They react to yield aqueous ammonium nitrate and solid lead(II) sulfide.

34. Copper metal is placed in a solution of silver nitrate. This produces aqueous copper(II) nitrate and silver metal.

35. The element nitrogen has two common isotopes: nitrogen-14 and nitrogen-16.

- State how these two types of atoms are similar.
- State how these two types of atoms are different.

Chemistry Answers

Section I. Multiple Choice

1. c
2. c
3. a
4. a
5. b
6. d
7. c
8. c
9. b
10. c
11. b
12. c
13. c
14. a
15. c

Section II. Skills

16. calcium bromide
17. gold(III) phosphate
18. dinitrogen tetroxide
19. ammonia
20. sulfuric acid
21. $\text{PbS}_{2(s)}$
22. $\text{CH}_{4(g)}$
23. $\text{SO}_{3(g)}$
24. $\text{HCl}_{(aq)}$
25. $\text{Fe}_3\text{N}_{2(s)}$

Section III. Response

26. 4, 1, 2
27. 1, 12, 8, 8
28. 4, 5, 4, 6
29. $\text{NaBr}_{(s)}$ formation
30. $\text{CO}_{2(g)} + \text{H}_2\text{O}_{(g)}$ hydrocarbon combustion
31. $\text{AlCl}_{3(aq)} + \text{Cu}_{(s)}$ single replacement
32. $\text{NaNO}_{3(aq)} + \text{PbI}_{2(s)}$ double replacement
33. $(\text{NH}_4)_2\text{S}_{(aq)} + \text{Pb}(\text{NO}_3)_2(aq) \rightarrow 2 \text{NH}_4\text{NO}_3(aq) + \text{PbS}_{(s)}$
34. $\text{Cu}_{(s)} + 2 \text{AgNO}_3(aq) \rightarrow \text{Cu}(\text{NO}_3)_2(aq) + 2 \text{Ag}_{(s)}$
35.
 - a) Both atoms have the same number of protons or atomic number.
 - b) One has 7 neutrons and the other has 8 neutrons. Their mass numbers are different.

Chemistry 10 Placement Test Data Booklet

Solubility of Some Common Ionic Compounds in Water at 25°C								
Ion	Group1 NH ₄ ⁺ H ₃ O ⁺ ,H ⁺	ClO ₃ ⁻ NO ₃ ⁻ ClO ₄ ⁻	CH ₃ COO ⁻	Cl ⁻ Br ⁻ I ⁻	SO ₄ ²⁻	S ²⁻	OH ⁻	PO ₄ ³⁻ SO ₃ ²⁻ CO ₃ ²⁻
Solubility greater than or equal to 0.1 mol/L (very soluble)	all	all	most	most	most	Group1 Group2 NH ₄ ⁺	Group1 NH ₄ ⁺ Sr ²⁺ Ba ²⁺ Tl ⁺	Group1 NH ₄ ⁺
Solubility less than 0.1 mol/L (slightly soluble)	none	none	Ag ⁺ Hg ⁺	Ag ⁺ Pb ²⁺ Hg ⁺ Cu ⁺ Tl ⁺	Ca ²⁺ Sr ²⁺ Ba ²⁺ Ra ²⁺ Pb ²⁺ Ag ⁺	most	most	most

1	2	3	4	5	6	7	8	9
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1 1.01 1+, 1- H hydrogen	3 6.94 1+ Li lithium	4 9.01 2+ Be beryllium
11 22.99 1+ Na sodium	12 24.31 2+ Mg magnesium	

acetate (ethanoate)	CH_3COO^-	chromate	CrO_4^{2-}	phosphate	PO_4^{3-}
ammonium	NH_4^+	dichromate	$\text{Cr}_2\text{O}_7^{2-}$	hydrogen phosphate	HPO_4^{2-}
benzoate	$\text{C}_6\text{H}_5\text{COO}^-$	cyanide	CN^-	dihydrogen phosphate	H_2PO_4^-
borate	BO_3^{3-}	hydroxide	OH^-	silicate	SiO_3^{2-}
carbide	C_2^{2-}	iodate	IO_3^-	sulfate	SO_4^{2-}
carbonate	CO_3^{2-}	nitrate	NO_3^-	hydrogen sulfate	HSO_4^-
hydrogen carbonate	HCO_3^-	nitrite	NO_2^-	sulfite	SO_3^{2-}
perchlorate	ClO_4^-	oxalate	$\text{O}^{2-}\text{C}^{2-}\text{C}^{2-}\text{O}^{2-}$	hydrogen sulfite	HSO_3^-
chlorate	ClO_3^-	hydrogen oxalate	$\text{HO}^{2-}\text{C}^{2-}\text{C}^{2-}\text{O}^{2-}$	hydrogen sulfide	HS^-
chlorite	ClO_2^-	permanganate	MnO_4^-	thiocyanate	SCN^-
hypochlorite	OCl^- or ClO^-	peroxide	O_2^{2-}	thiosulfate	$\text{S}_2\text{O}_3^{2-}$
		persulfide	S_2^{2-}		

19 39.10 1+ 0.8 K potassium	20 40.08 2+ 1.0 Ca calcium	21 44.96 3+ 1.4 Sc scandium	22 47.87 4+, 3+ 1.5 Ti titanium	23 50.94 5+, 4+ 1.6 V vanadium	24 52.00 3+, 2+ 1.7 Cr chromium	25 54.94 2+, 4+ 1.6 Mn manganese	26 55.85 3+, 2+ 1.8 Fe iron	27 58.93 2+, 3+ 1.9 Co cobalt
37 85.47 1+ 0.8 Rb rubidium	38 87.62 2+ 1.0 Sr strontium	39 88.91 3+ 1.2 Y yttrium	40 91.22 4+ 1.3 Zr zirconium	41 92.91 5+, 3+ 1.6 Nb niobium	42 95.94 6+ 2.2 Mo molybdenum	43 (98) 7+ 2.1 Tc technetium	44 101.07 3+ 2.2 Ru ruthenium	45 102.91 3+ 2.3 Rh rhodium
55 132.91 1+ 0.8 Cs cesium	56 137.33 2+ 0.9 Ba barium	57 138.91 3+ 1.1 La lanthanum	72 178.49 4+ 1.3 Hf hafnium	73 180.95 5+ 1.5 Ta tantalum	74 183.84 6+ 1.7 W tungsten	75 186.21 7+ 1.9 Re rhenium	76 190.23 4+ 2.2 Os osmium	77 192.22 4+ 2.2 Ir iridium
87 (223) 1+ 0.7 Fr francium	88 (226) 2+ 0.9 Ra radium	89 (227) 3+ 1.1 Ac actinium	104 (261) 4+ Rf rutherfordium	105 (262) Db dubnium	106 (266) Sg seaborgium	107 (264) Bh bohrium	108 (277) Hs hassium	109 (268) Mt meitnerium

—lanthanide and actinide series begin

References

Lide, D.R. 2005. *CRC Handbook of Chemistry and Physics*. 86th ed. Boca Raton: CRC Press.

Speight, James G. 2005. *Lange's Handbook of Chemistry*. 16th ed. New York: McGraw-Hill, Inc.

IUPAC commission on atomic weights and isotopic abundances. 2002. <http://www.chem.qmw.ac.uk/iupac/AtWV/index.html>.

58 140.12 3+ 1.1 Ce cerium	59 140.91 3+ 1.1 Pr praseodymium	60 144.24 3+ 1.1 Nd neodymium	61 (145) 3+ — Pm promethium	62 150.36 3+, 2+ 1.2 Sm samarium
90 232.04 4+ 1.3 Th thorium	91 231.04 5+, 4+ 1.5 Pa protactinium	92 238.03 6+, 4+ 1.7 U uranium	93 (237) 5+ 1.3 Np neptunium	94 (244) 4+, 6+ 1.3 Pu plutonium

10	11	12	13	14	15	16	17	18
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Legend for Elements

 Metallic solids	 Gases
 Non-metallic solids	 Liquids

Note: The legend denotes the physical state of the elements at exactly 101.325 kPa and 298.15 K.

Key

Atomic number	26	55.85
Electronegativity	1.8	3+, 2+
Symbol	Fe	
Name	iron	

Atomic molar mass (g/mol)*
Most stable ion charges

* Based on $^{12}_6\text{C}$
() Indicates mass of the most stable isotope

										2 4.00 He helium	
		5 10.81 B boron	6 12.01 C carbon	7 14.01 N nitrogen	8 16.00 O oxygen	9 19.00 F fluorine	10 20.18 Ne neon				
		13 26.98 Al aluminium	14 28.09 Si silicon	15 30.97 P phosphorus	16 32.07 S sulfur	17 35.45 Cl chlorine	18 39.95 Ar argon				
28 58.69 1.9 Ni nickel	29 63.55 1.9 Cu copper	30 65.41 1.7 Zn zinc	31 69.72 1.8 Ga gallium	32 72.64 2.0 Ge germanium	33 74.92 2.2 As arsenic	34 78.96 2.6 Se selenium	35 79.90 3.0 Br bromine	36 83.80 — Kr krypton			
46 106.42 2.2 Pd palladium	47 107.87 1.9 Ag silver	48 112.41 1.7 Cd cadmium	49 114.82 1.8 In indium	50 118.71 2.0 Sn tin	51 121.76 2.1 Sb antimony	52 127.60 2.1 Te tellurium	53 126.90 2.7 I iodine	54 131.29 2.6 Xe xenon			
78 195.08 2.2 Pt platinum	79 196.97 2.4 Au gold	80 200.59 1.9 Hg mercury	81 204.38 1.8 Tl thallium	82 207.2* 1.8 Pb lead	83 208.98 1.9 Bi bismuth	84 (209) 2.0 Po polonium	85 (210) 2.2 At astatine	86 (222) — Rn radon			
110 (271) Ds darmstadtium	111 (272) Rg roentgenium										* The isotopic mix of naturally occurring lead is more variable than other elements, preventing precision to greater than tenths of a gram per mole.

63 151.96 — Eu europium	64 157.25 1.2 Gd gadolinium	65 158.93 — Tb terbium	66 162.50 1.2 Dy dysprosium	67 164.93 1.2 Ho holmium	68 167.26 1.2 Er erbium	69 168.93 1.3 Tm thulium	70 173.04 — Yb ytterbium	71 174.97 1.0 Lu lutetium			
95 (243) — Am americium	96 (247) — Cm curium	97 (247) — Bk berkelium	98 (251) — Cf californium	99 (252) — Es einsteinium	100 (257) — Fm fermium	101 (258) — Md mendelevium	102 (259) — No nobelium	103 (262) — Lr lawrencium			